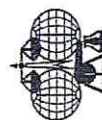


IUPAC Periodic Table of the Elements

18	2	He	helium	4.0026
17	9	F	fluorine	18.998
16	8	O	oxygen	15.999
15	7	N	nitrogen	14.007
14	6	C	carbon	12.011
13	5	B	boron	10.811
12	4	Be	beryllium	9.0122
11	3	Li	lithium	6.941
10	2	H	hydrogen	1.008
9	1	H	hydrogen	1.008
8	1	H	hydrogen	1.008
7	1	H	hydrogen	1.008
6	1	H	hydrogen	1.008
5	1	H	hydrogen	1.008
4	1	H	hydrogen	1.008
3	1	H	hydrogen	1.008
2	1	H	hydrogen	1.008
1	1	H	hydrogen	1.008
118	117	Ts	tennessine	289.101
116	116	Lv	livermorium	293.101
115	115	Mc	moscovium	288.101
114	114	Fl	flerovium	289.101
113	113	Nh	nihonium	284.101
112	112	Cn	copernicium	285.101
111	111	Rg	roentgenium	282.101
110	110	Ds	darmstadtium	281.101
109	109	Mt	meitnerium	288.101
108	108	Hs	hassium	277.101
107	107	Bh	bohrium	264.101
106	106	Sg	seaborgium	266.101
105	105	Db	dubnium	261.101
104	104	Rf	rutherfordium	261.101
103	103	La	lanthanum	138.905
102	102	Yb	ytterbium	173.054
101	101	Lu	lutetium	174.967
100	100	Fm	fermium	150.107
99	99	Es	eesmanium	148.107
98	98	Cf	californium	151.083
97	97	Bk	berkelium	147.077
96	96	Cm	curium	147.077
95	95	Am	americium	151.063
94	94	Pu	plutonium	150.947
93	93	Np	neptunium	150.074
92	92	U	uranium	238.029
91	91	Pa	protactinium	231.036
90	90	Th	thorium	232.038
89	89	Ac	actinium	227.028
88	88	Ra	radium	226.025
87	87	Fr	francium	223.019
86	86	Rn	radon	222.018
85	85	At	astatine	210.084
84	84	Po	polonium	209.083
83	83	Bi	bismuth	208.980
82	82	Pb	lead	207.2
81	81	Tl	thallium	204.383
80	80	Hg	mercury	200.59
79	79	Au	gold	196.967
78	78	Pt	platinum	195.084
77	77	Ir	iridium	192.222
76	76	Os	osmium	190.233
75	75	Re	rhenium	186.207
74	74	W	tungsten	183.84
73	73	Ta	tantalum	180.948
72	72	Hf	hafnium	178.49
71	71	La	lanthanum	138.905
70	70	Yb	ytterbium	173.054
69	69	Tm	thulium	168.934
68	68	Er	erbium	167.259
67	67	Ho	holmium	164.930
66	66	Dy	dysprosium	162.500
65	65	Tb	terbium	158.925
64	64	Gd	gadolinium	157.254
63	63	Eu	euporium	151.964
62	62	Sm	samarium	150.36
61	61	Pm	promethium	144.913
60	60	Nd	neodymium	144.242
59	59	Pr	praseodymium	140.908
58	58	Ce	cerium	140.127
57	57	La	lanthanum	138.905
56	56	Ba	barium	137.327
55	55	Cs	caesium	132.905
54	54	Xe	xenon	131.29
53	53	I	iodine	126.905
52	52	Te	tellurium	127.603
51	51	Sb	antimony	121.757
50	50	Sn	tin	118.710
49	49	In	indium	114.818
48	48	Cd	cadmium	112.411
47	47	Ag	silver	107.868
46	46	Pd	palladium	106.421
45	45	Rh	rhodium	102.906
44	44	Ru	ruthenium	101.072
43	43	Tc	technetium	98.906
42	42	Mo	molybdenum	95.94
41	41	Nb	niobium	92.906
40	40	Zr	zirconium	91.224
39	39	Y	yttrium	88.906
38	38	Sr	strontium	87.62
37	37	Rb	rubidium	85.468
36	36	Kr	krypton	83.798
35	35	Br	bromine	79.904
34	34	Se	selenium	78.96
33	33	As	arsenic	74.922
32	32	Ge	germanium	72.630
31	31	Ga	gallium	69.723
30	30	Zn	zinc	65.38
29	29	Cu	copper	63.546
28	28	Ni	nickel	58.693
27	27	Co	cobalt	58.933
26	26	Fe	iron	55.845
25	25	Mn	manganese	54.938
24	24	Cr	chromium	51.996
23	23	V	vanadium	50.942
22	22	Ti	titanium	47.867
21	21	Sc	scandium	44.956
20	20	Ca	calcium	40.078
19	19	K	potassium	39.098
18	18	Ar	argon	39.948
17	17	Cl	chlorine	35.453
16	16	S	sulfur	32.06
15	15	P	phosphorus	30.974
14	14	Si	silicon	28.086
13	13	Al	aluminum	26.982
12	12	Mg	magnesium	24.305
11	11	Na	sodium	22.990
10	10	Ne	neon	20.180
9	9	F	fluorine	18.998
8	8	O	oxygen	15.999
7	7	N	nitrogen	14.007
6	6	C	carbon	12.011
5	5	B	boron	10.811
4	4	Be	beryllium	9.0122
3	3	Li	lithium	6.941
2	2	H	hydrogen	1.008
1	1	H	hydrogen	1.008

Key:
atomic number
symbol
name
conventional atomic weight
standard atomic weight



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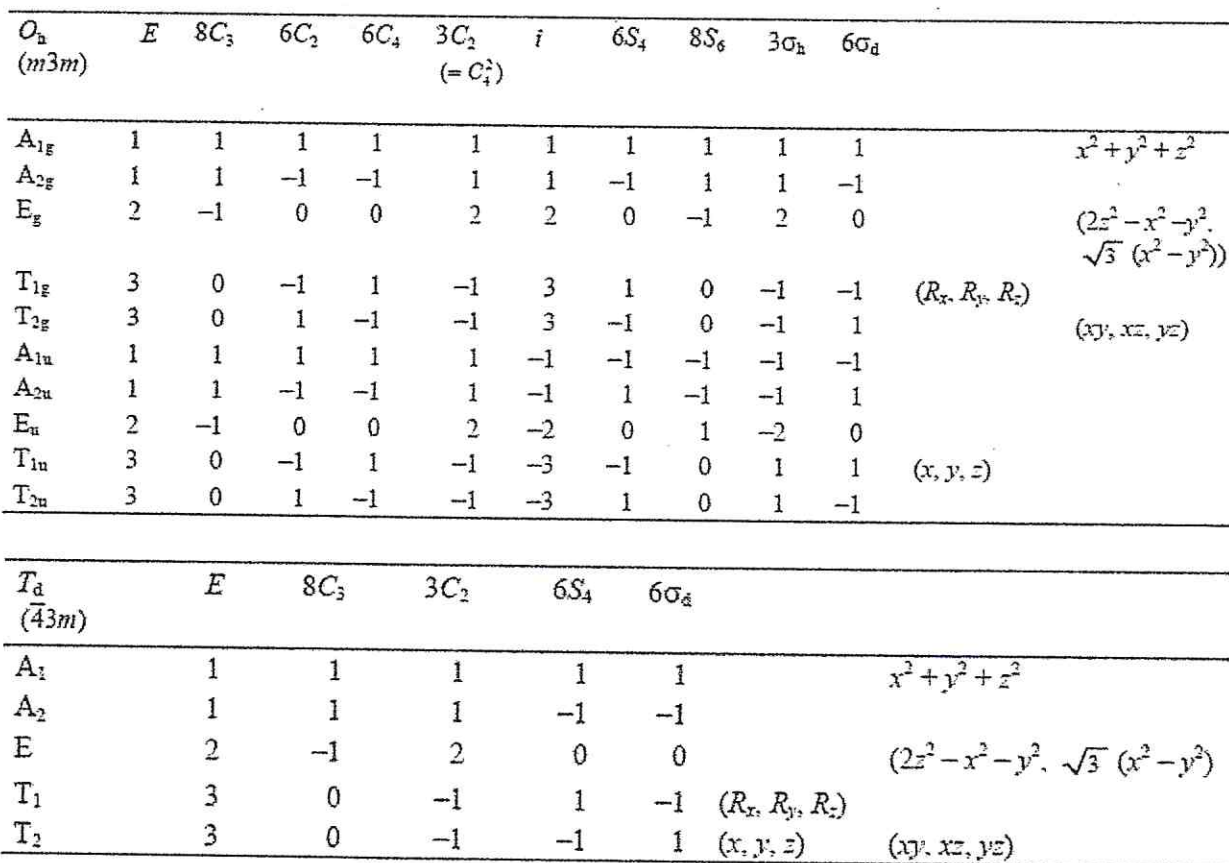
For notes and updates to this table, see www.iupac.org. This version is dated 28 November 2016.
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Material Suplementar

Fórmulas, constantes, unidades e conversões		
$\left(\frac{\partial \Delta G}{\partial T}\right)_p = -nF \left(\frac{\partial E}{\partial T}\right)_p$	$E = E^0 - \frac{RT}{nF} \ln \frac{\prod_i [a_{\text{produtos}}]_i^{\delta_i}}{\prod_i [a_{\text{reagentes}}]_i^{\delta_i}}$	1 pm = 10 ⁻¹² m
$\Delta S = nF \left(\frac{\partial E}{\partial T}\right)_p$	R = 8,314 J mol ⁻¹ K ⁻¹	1 Å = 10 ⁻¹⁰ m
$\Delta G = \Delta H - T\Delta S$	F = 96.485 C mol ⁻¹	1 nm = 10 ⁻⁹ m
$dS = \frac{dq_{\text{rev}}}{T}$	$a_i = c_i \gamma_i$	0 °C = 273 K
$\Delta G = -nFE$	$\gamma_{\pm}^{\nu} = \gamma_+^{\nu_+} \gamma_-^{\nu_-}$	1 eV = 1,60 x 10 ⁻¹⁹ J
$k = \frac{a_0 - a}{t}$	$\log \gamma_{\pm} = -A Z_+ Z_- \sqrt{I}$	1 J = 1 kg m ² s ⁻²
$\log(a - x) = \log a - \frac{kt}{2,303}$	$k = \frac{1}{t} \ln \left(\frac{a_0}{a_0 - x} \right)$	J = VC
$t_{1/2} = \frac{\ln 2}{k}$	$\Delta S_T = \Delta S_{\text{sistema}} + \Delta S_{\text{viz}} > 0$	c = 3,0 x 10 ⁸ m s ⁻¹
$k = \frac{1}{(a_0 - b_0)t} \ln \left[\frac{b_0(a_0 - x)}{a_0(b_0 - x)} \right]$	$A = \epsilon b C$	ln(x) = 2,303 log(x)
$\beta = \frac{B_{\text{complexo}}}{B_{\text{ion livre}}}$	$A_{\text{Total}} = \epsilon_x b C_x + \epsilon_y b C_y$	K _w = 1,0 x 10 ⁻¹⁴ (25 °C)

Potenciais Padrão de Eletrodos*

Reação	E ⁰ a 25 °C, V
Cl ₂ (g) + 2e ⁻ ⇌ 2Cl ⁻	+1,359
O ₂ (g) + 4H ⁺ + 4e ⁻ ⇌ 2H ₂ O	+1,229
Br ₂ (aq) + 2e ⁻ ⇌ 2Br ⁻	+1,087
Br ₂ (l) + 2e ⁻ ⇌ 2Br ⁻	+1,065
Ag ⁺ + e ⁻ ⇌ Ag(s)	+ 0,799
Fe ³⁺ + e ⁻ ⇌ Fe ²⁺	+ 0,771
I ₂ + 2e ⁻ ⇌ 2I ⁻	+ 0,536
Cu ²⁺ + 2e ⁻ ⇌ Cu(s)	+ 0,337
UO ₂ ²⁺ + 4H ⁺ + 2e ⁻ ⇌ U ⁴⁺ + 2H ₂ O	+ 0,334
Hg ₂ Cl ₂ (s) + 2e ⁻ ⇌ 2Hg(l) + 2Cl ⁻	+ 0,268
AgCl(s) + e ⁻ ⇌ Ag(s) + Cl ⁻	+ 0,222
Ag(S ₂ O ₃) ₃ ³⁻ + e ⁻ ⇌ Ag(s) + 2S ₂ O ₃ ²⁻	+ 0,017
2H ⁺ + 2e ⁻ ⇌ H ₂ (g)	0,000
AgI(s) + e ⁻ ⇌ Ag(s) + I ⁻	- 0,151
PbSO ₄ + 2e ⁻ ⇌ Pb(s) + SO ₄ ²⁻	- 0,350
Cd ²⁺ + 2e ⁻ ⇌ Cd(s)	- 0,403
Zn ²⁺ + 2e ⁻ ⇌ Zn(s)	- 0,763
CrSO ₄ (s) + 2e ⁻ ⇌ Cr(s) + SO ₄ ²⁻ (aq)	-0,4



C_{2v} ($2mm$)	E	C_2	$\sigma_v(xz)$	$\sigma'_v(yz)$		
A_1	1	1	1	1	z	x^2, y^2, z^2
A_2	1	1	-1	-1	R_z	xy
B_1	1	-1	1	-1	x, R_y	xz
B_2	1	-1	-1	1	y, R_x	yz

C_{3v} ($3m$)	E	$2C_3$	$3\sigma_v$		
A_1	1	1	1	z	$x^2 + y^2, z^2$
A_2	1	1	-1	R_z	
E	2	-1	0	$(x, y)(R_x, R_y)$	$(x^2 - y^2, 2xy)(xz, yz)$

C_{4v} ($4mm$)	E	$2C_4$	C_2	$2\sigma_v$	$2\sigma_d$		
A_1	1	1	1	1	1	z	$x^2 + y^2, z^2$
A_2	1	1	1	-1	-1	R_z	
B_1	1	-1	1	1	-1		$x^2 - y^2$
B_2	1	-1	1	-1	1		xy
E	2	0	-2	0	0	$(x, y)(R_x, R_y)$	(xz, yz)

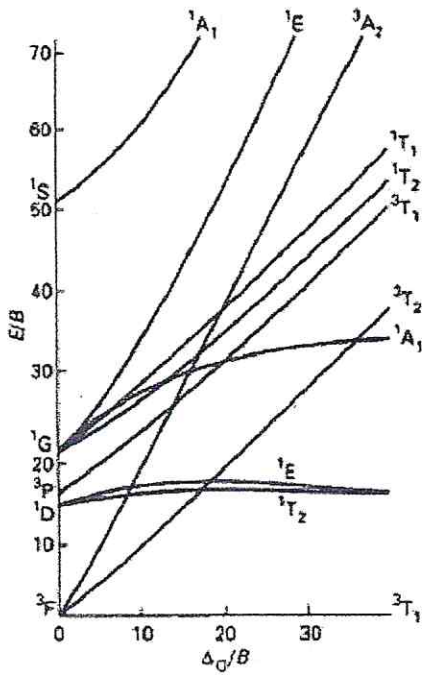
C_{2h} ($2/m$)	E	C_2	I	σ_h		
A_g	1	1	1	1	R_z	x^2, y^2, z^2, xy
B_g	1	-1	1	-1	R_x, R_y	xz, yz
A_u	1	1	-1	-1	z	
B_u	1	-1	-1	1	x, y	

C_{3h} ($\bar{6}$)	E	C_3	C_3^2	σ_h	S_3	S_3^5	$\varepsilon = \exp(2\pi i/3)$	
A'	1	1	1	1	1	1	R_z	$x^2 + y^2, z^2$
E'	$\begin{Bmatrix} 1 & \varepsilon & \varepsilon^* \\ 1 & \varepsilon^* & \varepsilon \end{Bmatrix}$	$\begin{Bmatrix} \varepsilon & 1 \\ \varepsilon^* & 1 \end{Bmatrix}$	$\begin{Bmatrix} \varepsilon^* & 1 \\ \varepsilon & 1 \end{Bmatrix}$	$\begin{Bmatrix} 1 & \varepsilon \\ 1 & \varepsilon^* \end{Bmatrix}$	$\begin{Bmatrix} \varepsilon & \varepsilon^* \\ \varepsilon^* & \varepsilon \end{Bmatrix}$	$\begin{Bmatrix} \varepsilon^* & \varepsilon \\ \varepsilon & \varepsilon^* \end{Bmatrix}$	(x, y)	$(x^2 - y^2, 2xy)$
A''	1	1	1	-1	-1	-1	z	
E''	$\begin{Bmatrix} 1 & \varepsilon & \varepsilon^* \\ 1 & \varepsilon^* & \varepsilon \end{Bmatrix}$	$\begin{Bmatrix} \varepsilon & 1 \\ \varepsilon^* & 1 \end{Bmatrix}$	$\begin{Bmatrix} \varepsilon^* & 1 \\ \varepsilon & 1 \end{Bmatrix}$	$\begin{Bmatrix} 1 & \varepsilon \\ 1 & \varepsilon^* \end{Bmatrix}$	$\begin{Bmatrix} \varepsilon & \varepsilon^* \\ \varepsilon^* & \varepsilon \end{Bmatrix}$	$\begin{Bmatrix} \varepsilon^* & \varepsilon \\ \varepsilon & \varepsilon^* \end{Bmatrix}$	(R_x, R_y)	(xz, yz)

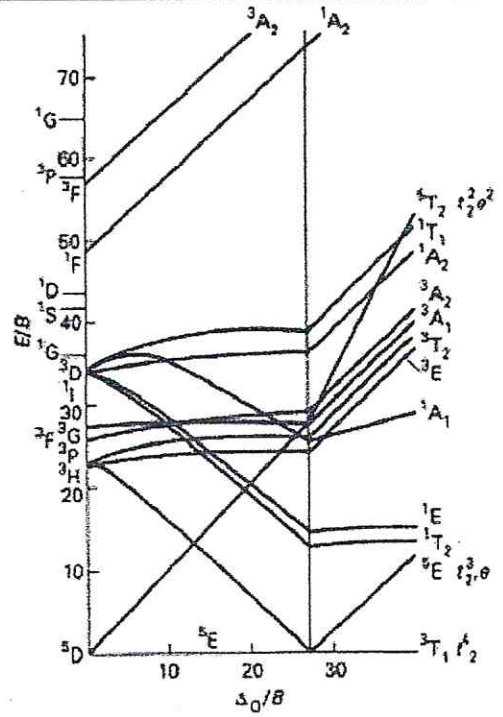
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Diagramas de Tanabe-Sugano

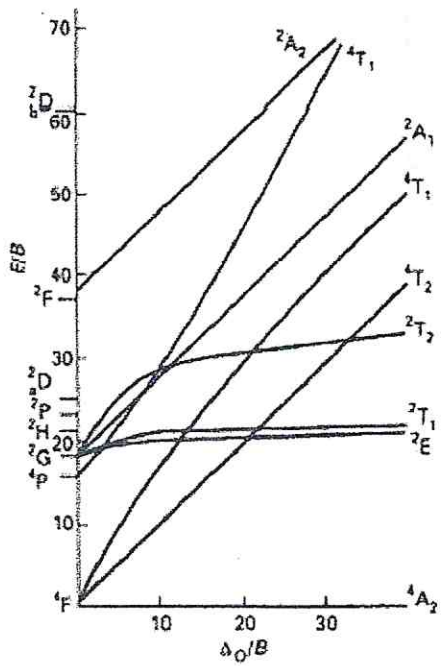
1. d^2 with $C = 4.42B$



3. d^* with $C = 4.61B$



2. d^2 with $C = 4.5B$



4. d^5 with $C = 4.477B$

